

05/28/19

Low Level Alkalinity Manual Titration Analysis Log

Analyst AK

Ident	Amount(ml)	Initial pH	pH(4.3-4.7)	Final pH	0.02N H2SO4				Alkalinity	Date/Time
					A	B	C	D=2(B-A)-(C-A)		
					Initial (ml)	ml Reading at pH(4.3-4.7)	Final(ml)	Difference(ml)		
LB103026BLW	100	4.56	4.56	4.26	0	0	0.04	-0.04	-0.4	5/24/2019 13:20
LB103026BSW	100	10.21	4.48	4.18	0	5.4	5.56	5.24	52.4	5/24/2019 13:25
K3032-08	100	5.26	4.58	4.28	0	0.08	0.14	0.02	0.2	5/24/2019 13:31
K3032-08DUP	100	5.28	4.56	4.22	0	0.1	0.14	0.06	0.6	5/24/2019 13:37
								0	#DIV/0!	5/24/2019 13:37
								0	#DIV/0!	5/24/2019 13:37
								0	#DIV/0!	5/24/2019 13:37
								0	#DIV/0!	5/24/2019 13:37
								0	#DIV/0!	5/24/2019 13:37
								0	#DIV/0!	5/24/2019 13:37
								0	#DIV/0!	5/24/2019 13:37
								0	#DIV/0!	5/24/2019 13:37
								0	#DIV/0!	5/24/2019 13:37
								0	#DIV/0!	5/24/2019 13:37
								0	#DIV/0!	5/24/2019 13:37
								0	#DIV/0!	5/24/2019 13:37
								0	#DIV/0!	5/24/2019 13:37
								0	#DIV/0!	5/24/2019 13:37
Note - Final pH has to be exactly 0.3 less than that observed at pH(4.3-4.7)										

Note - Final pH has to be exactly 0.3 less than that observed at pH(4.3-4.7)

Alkalinity = $D \cdot N \cdot 50000$ /ml of sample

where N Is Normality of sulfuric acid = 0.02