

Low Level Alkalinity Manual Titration Analysis Log

Analyst AK

Ident	Amount(ml)	Initial pH	pH(4.3-4.7)	Final pH	0.02N H2SO4				Alkalinity	Date/Time
					A	B	C	D=2(B-A)-(C-A)		
					Initial (ml)	ml Reading at pH(4.3-4.7)	Final(ml)	Difference(ml)		
LB102371BLW	100	4.34	4.34	4.04	0	0	0.06	-0.06	-0.6	4/24/2019 10:20
LB102371BSW	100	10.25	4.58	4.28	0	5.46	5.6	5.32	53.2	4/24/2019 10:30
K2487-06	100	4.80	4.50	4.20	0	0.06	0.1	0.02	0.2	4/24/2019 10:37
K2487-06DUP	100	4.81	4.52	4.22	0	0.06	0.1	0.02	0.2	4/24/2019 10:50
								0	#DIV/0!	4/24/2019 10:50
								0	#DIV/0!	4/24/2019 10:50
								0	#DIV/0!	4/24/2019 10:50
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								0	#DIV/0!	4/24/2019 10:50

Note - Final pH has to be exactly 0.3 less than that observed at pH(4.3-4.7)

Alkalinity = $D \cdot N \cdot 50000$ /ml of sample

where N Is Normality of sulfuric acid = 0.02