

AD 07/01/19

Low Level Alkalinity Manual Titration Analysis Log

Analyst AK

Ident	Amount(ml)	Initial pH	pH(4.3-4.7)	Final pH	0.02N H2SO4				Alkalinity	Date/Time
					A	B	C	D=2(B-A)-(C-A)		
					Initial (ml)	ml Reading at pH(4.3-4.7)	Final(ml)	Difference(ml)		
LB103720BLW	100	4.48	4.48		0	0	0.04	-0.04	-0.4	6/27/2019 16:20
LB103720BSW	100	9.98	4.60		0	5.38	5.8	4.96	49.6	6/27/2019 16:24
K3521-01	100	4.33	4.33		0	0	0.08	-0.08	-0.8	6/27/2019 16:29
K3521-03	100	5.58	4.66		0	0.32	0.06	0.58	5.8	6/27/2019 16:35
K3521-36	100	5.30	4.38		0	0.68	0.1	1.26	12.6	6/27/2019 16:41
K3521-37	100	5.41	4.42		0	1	0.08	1.92	19.2	6/27/2019 16:45
K3556-02	100	4.88	4.58		0	0.06	0.04	0.08	0.8	6/27/2019 16:50
K3556-02DUP	100	4.89	4.57		0	0.06	0.04	0.08	0.8	6/27/2019 16:56
								0	#DIV/0!	6/27/2019 16:56
								0	#DIV/0!	6/27/2019 16:56
								0	#DIV/0!	6/27/2019 16:56
								0	#DIV/0!	6/27/2019 16:56
								0	#DIV/0!	6/27/2019 16:56

Note - Final pH has to be exactly 0.3 less than that observed at pH(4.3-4.7)

Alkalinity = $D \cdot N \cdot 50000$ / ml of sample

where N is Normality of sulfuric acid = 0.02